Slide 73 answers were correct.  
  
All redox equations are technically reversible (which is the whole idea of electrolysis) but the equilibrium constant usually heavily favours one side of the reaction. e.g. when you hook up Ag+/Ag(s) half cell to Cu2+/Cu(s) half-cell, the reaction heavily favours Ag+ --> Ag(s) and Cu(s) --> Cu2+, not the other way around. That is why we often use single arrows when talking about a particular redox reaction.